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$\text{pH} = 4.76 + \log$
 $(4.04/5.96) = 4.59.$

Thus, a change in pH of 0.19 pH units will be observed. The

buffering capacity =
(change in OH-) /
(change in pH) = 1

mM/0.19 pH
units = 5.26 mmol/pH

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acid = 7.96 mmol and
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