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Concepts: Terms in this

set (57) Dalton. The

atom is a tiny,

indestructible particle

with no internal

structure. Thomson.

The atom is a sphere of

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positive electrical  
charge with electrons  
embedded in the  
sphere.

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Learn You will compare

the wave and particle

models of light. You

will describe how the

frequency of light

emitted by an atom is

a unique characteristic

of that atom. You will

compare and contrast

the Bohr and quantum

mechanical models of

the atom. You will

express the

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arrangements of  
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Chapter 6 - periodic  
table & trends

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- moles (handouts)  
Chapter 11 - reactions  
(handouts) Chapter 12  
...

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Atoms Models of the  
Atom Rutherford used  
existing ideas about  
the atom and proposed

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an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Chapter 5 - Electrons in Atoms. Chapter 5 - Electrons in Atoms.

Jennie L. Borders.

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## Chapter 5

Section 5.1 – Models of  
the Atom. The

Rutherford's model of  
the atom did not

explain how an atom  
can emit light or the  
chemical properties of  
an atom. Plum Pudding  
Model Rutherford's  
Model. The Bohr Model.

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Terms in this set (82)

Rutherford's model-

also called nuclear

model-did not explain

how the electrons were

arranged

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Chapter 5 - Electrons

in Atoms Section 5.1 -

Models of the Atom

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The Rutherford's model of the atom did not explain how an atom can emit light or the chemical properties of an atom. Plum Pudding Model Rutherford's Model

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pronouncement ...

Atoms Vocabulary

**Chapter 5 Electrons**

**In Atoms Work**

### **Answers**

...are the way electrons are arranged in various orbitals around the nuclei of atoms. Three rules tell us how:

Aufbau principle - electrons enter the lowest energy first.

This causes difficulties because of the overlap of orbitals of different energies - follow the



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diagram! Pauli  
Exclusion Principle - at  
most 2 electrons per  
orbital ...

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How many electrons  
can each p orbital  
hold? Chapter 5:  
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DRAFT. 10th - 11th  
grade. 60 times.  
Chemistry. 77%  
average accuracy. 2  
years ago,

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grade .

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Atoms. Educators. AY  
IB Chapter Questions.  
02:25. Problem 1  
Objects get their colors  
from reflecting only

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certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of  $4.90 \times 10^{-7} \text{ m}$ . What is the frequency of the light? ...

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Electrons in successive atoms on the periodic table tend to fill low-

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energy orbitals first. Thus, many students find it confusing that, for example, the 5p orbitals fill immediately after the 4d, and immediately before the 6s. The filling order is based on observed experimental results, and has been confirmed by theoretical calculations.

**Electronic Structure  
of Atoms | CHEM**

*Page 20/24*

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### **1305 Introductory ...**

Electrons are

extremely small. The

mass of an electron is

only about  $1/2000$  the

mass of a proton or

neutron, so electrons

contribute virtually

nothing to the total

mass of an atom.

Electrons have an ...

### **4.4: The Properties**

**of Protons,**

**Neutrons, and**

**Electrons ...**

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*Page 21/24*

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the wave and particle

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the atom. You will

express the

arrangements of

electrons in atoms

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through orbital

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### **Chapter 5: Electrons in Atoms -**

#### **Neshaminy School District**

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Although the speed of all electromagnetic waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the

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previous page,  
wavelength and  
frequency are inversely  
related; in other words,  
as one quantity  
increases, the other  
decreases.

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