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CHAPTER 12 Assessment 36. a. ... Acceptable answers include the idea of writing a ratio using the coefficients of two substances from a balanced equation as the number of moles of each substance reacting or

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balanced equation, the ratio between the number of moles.... a reactant that is totally consumed during a chemical reaction.... the study of quantitative relationships between the amounts of....

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Test. PLAY. Match. Gravity. Created by. MEG_RIST. Overview of Chemistry 1 Honors Chapter 12: Stoichiometry. Terms in this set (21) Stoichiometry. The calculation of quantities in chemical reactions is a subject of chemistry. Mole ratio. A conversion factor derived from coefficients of a balanced chemical equation interpreted in terms of moles.

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. ____ 1.

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The first step in most stoichiometry problems is to ____ . a. add the coefficients of the reagents c. convert given quantities to volumes b. convert given quantities to moles d. convert given quantities to masses ____ 2.

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e. 12 g 7. How many grams of nitric acid, HNO_3 , can be prepared from the reaction of 138 g of NO_2 with 54.0 g H_2O according to the equation below? $3\text{NO}_2 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$... Practice Test Ch3 Stoichiometry (page 3 of 3) 1. d It might be

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easiest to balance the equation with mostly whole numbers: $2\text{NH}_3 + \frac{1}{2}\text{O}_2 \rightarrow 2\text{NO} + 3\text{H}_2\text{O}$. The ...

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